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Throughout, Matsoukas focuses on topics that link tightly to other key areas of undergraduate chemical engineering, including separations, reactions, and capstone design. More than 300 end-of-chapter problems range from basic calculations to realistic environmental applications; these can be solved with any leading mathematical software.

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3. A solution is prepared by dissolving 20.0 g of ethylene glycol (1,2-ethanediol) in 1 kg of water. Estimate the boiling point and the freezing point of this solution. 4. A biopolymer has a molecular weight of 250,000 dalton. At 300 K, estimate the osmotic pressure of a solution that contains 1 g of this substance in 10 mL of water. 5.

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thermodynamics, we can predict the amount of energy needed to change a system from an equilibrium state to another. For example it will take about 75 kJ to change 1 kg of air at

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Solution. The gas constant that is most convenient to use is 8.314 J/K/mol. Therefore it is important to convert the kJ value of the heat of vaporization to J. \[\ln{\dfrac{350}{P_2}}=\dfrac{43990}{8.314} \left(\dfrac{1}{298}-\dfrac{1}{338}\right)\] Using basic algebra and the knowledge that \(e^{\ln(x)}=x\), we can solve for \(P_2\).

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A 0.6-m3 rigid tank is filled with saturated liquid water at 170°C. A valve at the bottom of the tank is now opened, and one-half of the total mass is withdrawn from the fcink in liquid form. Heat is transferred to water from a source of 210°C so that the temperature in the tank remains constant. Determine (a) the